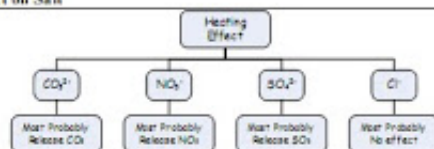


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Heating effect on Salt



Heating Effect on Carbonate Salt

Carbonate Salt	Equation of The Reaction
Potassium carbonate Sodium carbonate	Not decomposable
Calcium carbonate	$\text{CaCO}_3 \longrightarrow \text{CaO} + \text{CO}_2$
Magnesium carbonate	$\text{MgCO}_3 \longrightarrow \text{MgO} + \text{CO}_2$
Aluminium carbonate	$\text{Al}_2(\text{CO}_3)_3 \longrightarrow \text{Al}_2\text{O}_3 + 3\text{CO}_2$
Zinc carbonate	$\text{ZnCO}_3 \longrightarrow \text{ZnO} + \text{CO}_2$
Iron (III) carbonate	$\text{Fe}_2(\text{CO}_3)_3 \longrightarrow \text{Fe}_2\text{O}_3 + 3\text{CO}_2$
Lead (II) carbonate	$\text{PbCO}_3 \longrightarrow \text{PbO} + \text{CO}_2$
Copper (II) carbonate	$\text{CuCO}_3 \longrightarrow \text{CuO} + \text{CO}_2$
Mercury (II) carbonate	$2\text{HgCO}_3 \longrightarrow 2\text{Hg} + 2\text{CO}_2 + \text{O}_2$
Silver (I) carbonate	$2\text{Ag}_2\text{CO}_3 \longrightarrow 4\text{Ag} + 2\text{CO}_2 + \text{O}_2$
Ammonium carbonate	$(\text{NH}_4)_2\text{CO}_3 \longrightarrow \text{NH}_3 + \text{CO}_2 + \text{H}_2\text{O}$

Heating Effect on Nitrate Salt

Nitrate Salt	Equation of The Reaction
Potassium nitrate	$2\text{KNO}_3 \longrightarrow 2\text{KNO}_2 + \text{O}_2$
Sodium nitrate	$2\text{NaNO}_3 \longrightarrow 2\text{NaNO}_2 + \text{O}_2$
Calcium nitrate	$2\text{Ca}(\text{NO}_3)_2 \longrightarrow 2\text{CaO} + 4\text{NO}_2 + \text{O}_2$
Magnesium nitrate	$\text{Mg}(\text{NO}_3)_2 \longrightarrow 2\text{MgO} + 4\text{NO}_2 + \text{O}_2$
Aluminium nitrate	$4\text{Al}(\text{NO}_3)_3 \longrightarrow 2\text{Al}_2\text{O}_3 + 12\text{NO}_2 + 3\text{O}_2$
Zinc nitrate	$\text{Zn}(\text{NO}_3)_2 \longrightarrow 2\text{ZnO} + 4\text{NO}_2 + \text{O}_2$
Iron (III) nitrate	$4\text{Fe}(\text{NO}_3)_3 \longrightarrow 2\text{Fe}_2\text{O}_3 + 12\text{NO}_2 + 3\text{O}_2$
Lead (II) nitrate	$\text{Pb}(\text{NO}_3)_2 \longrightarrow 2\text{PbO} + 4\text{NO}_2 + \text{O}_2$
Copper (II) nitrate	$\text{Cu}(\text{NO}_3)_2 \longrightarrow 2\text{CuO} + 4\text{NO}_2 + \text{O}_2$
Mercury (II) nitrate	$\text{Hg}(\text{NO}_3)_2 \longrightarrow \text{Hg} + 2\text{NO}_2 + \text{O}_2$
Silver (I) nitrate	$2\text{AgNO}_3 \longrightarrow 2\text{Ag} + 2\text{NO}_2 + \text{O}_2$
Ammonium nitrate	$\text{NH}_4\text{NO}_3 \longrightarrow \text{N}_2\text{O} + 2\text{H}_2\text{O}$

[NOTES: Nitrogen dioxide, NO_2 is acidic gas and is brown in colour.]

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